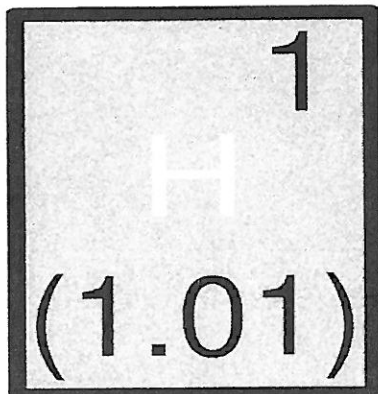


## The Atom

- \_\_\_\_\_ is made up of \_\_\_\_\_ (+) and \_\_\_\_\_ (no charge-neutral)
- Nucleus contains most of the atom's \_\_\_\_\_
- \_\_\_\_\_ is very \_\_\_\_\_ compared to rest of atom. Similar in size to a \_\_\_\_\_ on a football field.
- Nucleus has a \_\_\_\_\_ (+) charge (because of the \_\_\_\_\_) BUT the atom is usually \_\_\_\_\_. The \_\_\_\_\_ (-) are found around the nucleus & \_\_\_\_\_ out the protons (+).

The # of protons, electrons, & neutrons of an atom can be determined by using the periodic table.



- Atomic # = # of \_\_\_\_\_ and \_\_\_\_\_
- Mass # = # of protons \_\_\_\_\_ neutrons
- (Avg. Atomic mass is the decimal # on P.T. & mass # is just the average rounded)

### TO FIND THE # OF...

- Protons....it is the ATOMIC #
- Electrons....it is the ATOMIC #
- Neutrons....MASS # - ATOMIC #

Let's Try...

■ Cl  
P =        E =        N =

■ Mg  
P =        E =        N =

■ Ne  
P =        E =        N =