

Covalent Bonds:

- Formed when electrons are Shared between atoms.
- Always between two or more non metal atoms.
oxygen carbon dioxide
- Ex: O_2 CO_2

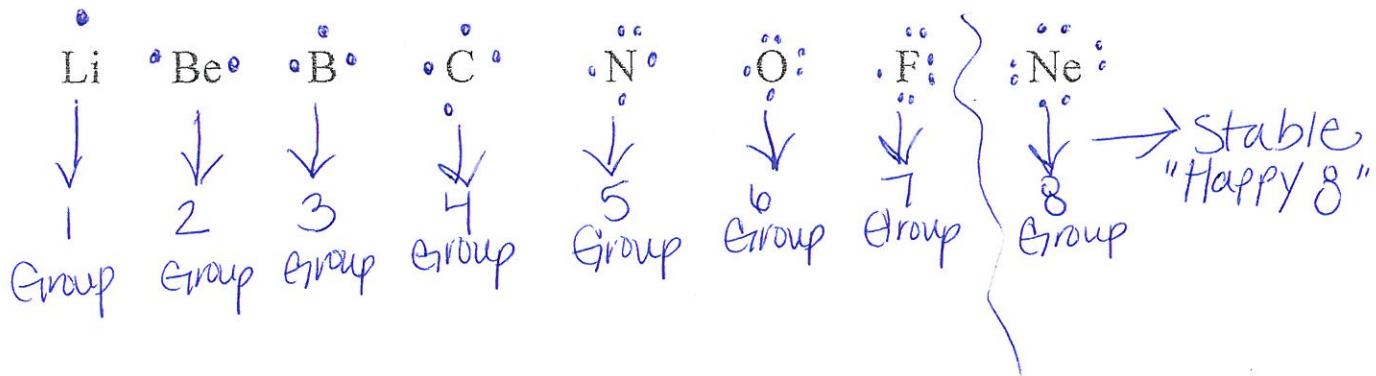
Metallic Bonds:

- Formed when electrons are easily kicked out from a large group of metal atoms creating metal ions. \rightarrow cations
- Occurs in metals atoms.
- The electrons are "freely floating" throughout the ions. "sea of electrons"
- The negative electrons holds the positive metal ions together (opposite charges attract)

Dot Diagrams:

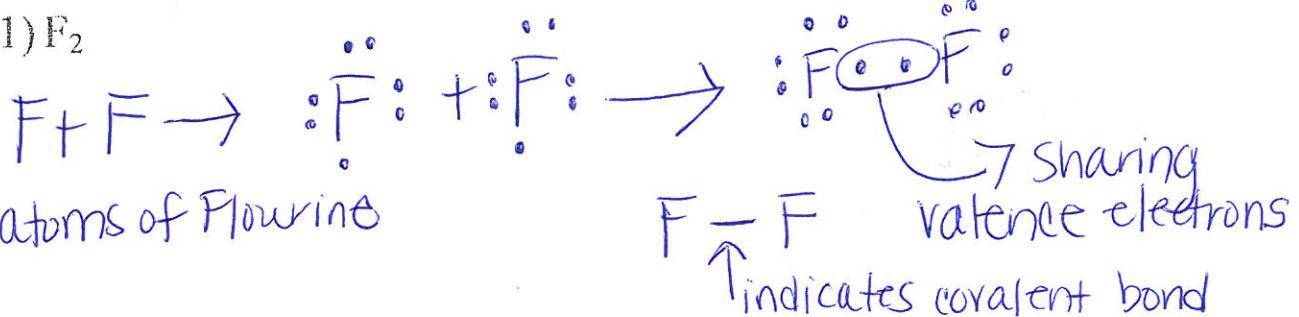
Represent the # of valence electrons an atom has and the bonds it can form.

How do you find the valence electrons?

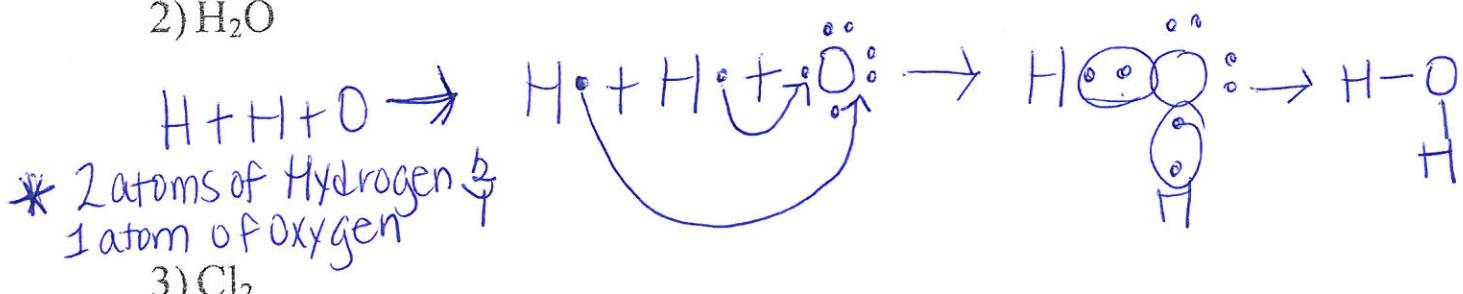


Covalent Bonds:

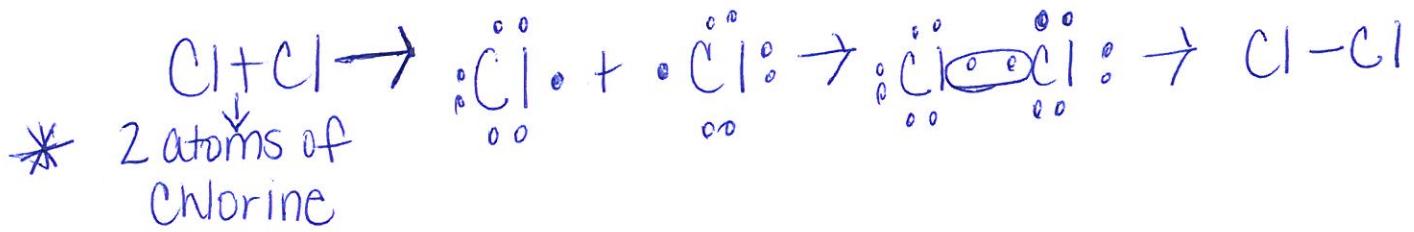
1) F₂



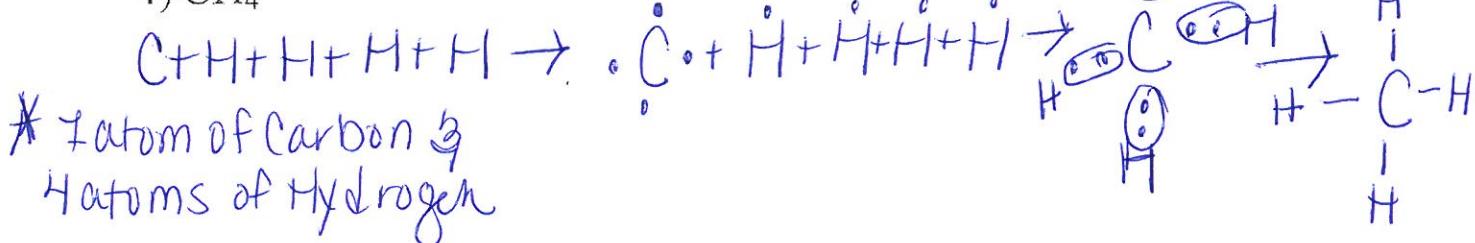
2) H₂O



3) Cl₂



4) CH₄



5) Br₂

